| | Application No. | Applicant(s) |
|---|---|---|
| ***Examiner-Initiated Interview Summary*** | 10/572,628 | AMOS ET AL. |
| | Examiner | Art Unit | |
| | M. Louisa Lao | 1621 | |

**All Participants:**

(1) *M. Louisa Lao*.

(2) *David Heiser*.

**Status of Application:** _____

(3) *Allen Sievert*.

(4) _____.

**Date of Interview:** *18 September 2007*          **Time:** *530pm*

**Type of Interview:**
☒ Telephonic
☐ Video Conference
☐ Personal (Copy given to: ☐ Applicant     ☐ Applicant's representative)

Exhibit Shown or Demonstrated:   ☐ Yes     ☐ No
If Yes, provide a brief description:              .

**Part I.**

Rejection(s) discussed:
*none*

Claims discussed:
*none*

Prior art documents discussed:
*none*

**Part II.**

SUBSTANCE OF INTERVIEW DESCRIBING THE GENERAL NATURE OF WHAT WAS DISCUSSED:
*Applicant suppled the step-wise procedure to arrive at the recitations for "atom percent" which entailed a mol to atom to atom percent calculation- which is summarized in the contnuation sheet..*

**Part III.**

☐ It is not necessary for applicant to provide a separate record of the substance of the interview, since the interview directly resulted in the allowance of the application. The examiner will provide a written summary of the substance of the interview in the Notice of Allowability.

☐ It is not necessary for applicant to provide a separate record of the substance of the interview, since the interview did not result in resolution of all issues. A brief summary by the examiner appears in Part II above.

_____          _____
(Examiner/SPE Signature)                    (Applicant/Applicant's/Representative Signature – if appropriate)

To calculate "atom percent"

assume 1000 gm sample into which is added 5% $Cr_2O_3$ (i.e. 50gm)
   step 1 - determine MW  where $Cr_2O_3$ MW= 152   $ZnCr_2O_4$ MW 233.4

   step 2- determine moles
      50/152 = 0.33 mol $Cr_2O_3$
      1000/233.4 = 4.28 mol $ZnCr_2O_4$

   step 3- determine atom equivalent of Cr in mol of components
      0.33 mol $Cr_2O_3$ is equivalent to 2 atoms Cr times 0.33  then = 0.66 atom equivalents Cr
      4.28 mol $ZnCr_2O_4$ is equiv to  2 Cr times 4.28  then = 8.56 atom equivalentst Cr

   step 4- determine atom percent Cr in 1050gms
      [0.66/ (8.56 + 0.66)] = 92.8 atom percent Cr

However, Applicants have not provided the atom percent calculatiions of the recited claims in reference to the working examples.